

- Enholm, E. J. & Jiang, S. (1992a). *Tetrahedron Lett.* pp. 313–316.
 Enholm, E. J. & Jiang, S. (1992b). *Tetrahedron Lett.* pp. 6069–6072.
 Sheldrick, G. M. (1976). *SHELX76*. Program for crystal structure determination. Univ. of Cambridge, England.
 Sheldrick, G. M. (1990). *SHELXTL-Plus*. Version 4.21/V. Siemens Analytical X-ray Instruments, Inc., Madison, Wisconsin, USA.
 Stewart, R. F., Davidson, E. R. & Simpson, W. T. (1965). *J. Chem. Phys.* **42**, 3175–3187.
 Still, W. C., Kahn, M. & Mitra, A. (1978). *J. Org. Chem.* **43**, 2923–2924.

Acta Cryst. (1993). **C49**, 1237–1238

Structure Determination of a Monohydrated Piperazinium Hydrogenphosphate

D. RIOU, TH. LOISEAU AND G. FERÉY

Laboratoire des Fluorures - URA 449, Faculté des Sciences, Université du Maine, 72017 Le Mans CEDEX, France

(Received 10 April 1992; accepted 13 January 1993)

Abstract

Piperazinium hydrogenphosphate monohydrate, $(\text{N}_2\text{C}_4\text{H}_{12})\cdot\text{HPO}_4\cdot\text{H}_2\text{O}$, is composed of a network of isolated PO_4 tetrahedra and organic six-membered rings of piperazine. The two amino groups of the piperazine are protonated and the rings are stacked along the [010] direction. The stability of the framework is ensured by strong O—H linkages between the O atoms from the PO_4 tetrahedra and H atoms from the diamine and water molecules.

Comment

The AlPO_4 -*n* family has been studied extensively over a few years. These compounds possess great similarities to the known zeolites whose frameworks are characterized by their large porosities. This property is fundamental for their activity in catalysis or as molecular sieves. One route to the synthesis of microporous structures uses an organic amine as a template. However, Tapp, Milestone & Bibby (1988) noted that this type of preparation sometimes leads to a simple amine phosphate at low temperatures. This paper describes the synthesis and X-ray structure determination of the piperazinium hydrogenphosphate $(\text{N}_2\text{C}_4\text{H}_{12})\cdot\text{HPO}_4\cdot\text{H}_2\text{O}$.

The title compound consists of a framework of isolated PO_4 tetrahedra interleaved with piperazine and H_2O molecules. The two amino groups of each piperazine molecule are protonated giving the formula $(\text{N}_2\text{C}_4\text{H}_{12})^{2+}$.

The organic rings thus constituted are stacked perpendicularly to the [010] direction at approximately the $y = 0$ and $y = \frac{1}{2}$ levels, whereas the PO_4 tetrahedra and the H_2O molecules are at the $y = \frac{1}{4}$ and $y = \frac{3}{4}$ levels (Fig. 1). The P atoms are distributed at one site in a tetrahedral coordination with three P—O bond lengths around 1.52 Å and a fourth of 1.589 (3) Å. This difference has previously been observed in $(\text{C}_2\text{H}_{10}\text{N}_2)\cdot\text{Na}_2\cdot(\text{HPO}_4)_2\cdot 6\text{H}_2\text{O}$ (Averbuch-Pouchot, Durif & Guitel, 1987) and $(\text{C}_2\text{H}_{10}\text{N}_2)\cdot\text{HPO}_4$ (Averbuch-Pouchot & Durif, 1987); in each case it is characteristic of the existence of an $(\text{HPO}_4)^{2-}$ hydrogenphosphate species. No residual electronic density is detected in the vicinity of the O(4) atom in $(\text{N}_2\text{C}_4\text{H}_{12})\cdot\text{HPO}_4\cdot\text{H}_2\text{O}$, but to satisfy at the electroneutrality of the compound it is reasonable to locate a hydroxyl group at this site. Each O atom [except O(4)] of the PO_4 tetrahedra makes two strong O—H linkages with distances ranging from 1.72 (4) to 1.92 (6) Å; this network of hydrogen bonds ensures the stability of the framework.

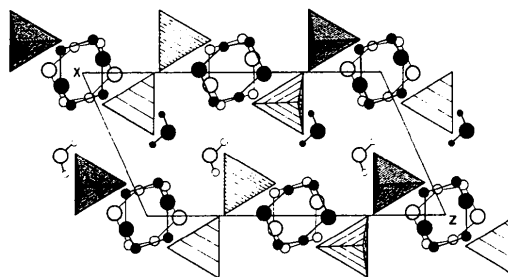


Fig. 1. Projection of $(\text{N}_2\text{C}_4\text{H}_{12})\cdot\text{HPO}_4\cdot\text{H}_2\text{O}$ along [010]; the piperazine rings and the H_2O molecules are drawn.

Experimental

Crystal data

$\text{C}_4\text{H}_{12}\text{N}_2^{2+}\cdot\text{HO}_4\text{P}^{2-}\cdot\text{H}_2\text{O}$
 $M_r = 202.16$
 Monoclinic
 $P2_1/c$
 $a = 6.4176$ (9) Å
 $b = 12.296$ (2) Å
 $c = 12.215$ (2) Å
 $\beta = 114.30$ (1)°
 $V = 878.5$ (3) Å³
 $Z = 4$
 $D_x = 1.529$ Mg m⁻³

Data collection

Stoe Siemens AED-2 diffractometer
 ω/θ scans
 Absorption correction: none
 4232 measured reflections
 4232 independent reflections
 891 observed reflections
 $[I > 3.0\sigma(I)]$

Mo $K\alpha$ radiation
 $\lambda = 0.71069$ Å
 Cell parameters from 34 reflections
 $\theta = 15$ – 16°
 $\mu = 0.294$ mm⁻¹
 $T = 293$ K
 Plate
 $0.228 \times 0.171 \times 0.027$ mm
 Colorless

$R_{\text{int}} = 0.0$
 $\theta_{\text{max}} = 27.5^\circ$
 $h = -8 \rightarrow 8$
 $k = 0 \rightarrow 15$
 $l = 0 \rightarrow 15$
 3 standard reflections
 frequency: 60 min
 intensity variation: 2.1%

Refinement

Refinement on F	$(\Delta/\sigma)_{\max} = 0.002$
Final $R = 0.039$	$\Delta\rho_{\max} = 0.34 \text{ e } \text{\AA}^{-3}$
$wR = 0.034$	$\Delta\rho_{\min} = -0.30 \text{ e } \text{\AA}^{-3}$
891 reflections	Atomic scattering factors
155 parameters	from <i>International Tables</i>
All H-atom parameters re-	for <i>X-ray Crystallogra-</i>
finied, $U(\text{H})$ with a com-	phy (1974, Vol. IV, Table
mon value	2.2B) (P, O) and Cromer
$w = 2.0811/[\sigma^2(F)$	& Mann (1968) (C, H, N)
$+ 0.0002F^2]$	

Data collection: Stoe *DIF4*. Cell refinement: Stoe *DIF4*. Data reduction: Stoe *REDU4*. Program(s) used to solve structure: *SHELXS86* (Sheldrick, 1985), option *TREF*. Program(s) used to refine structure: *SHELX76* (Sheldrick, 1976).

Table 1. Fractional atomic coordinates and equivalent isotropic thermal parameters (\AA^2)

$$U_{\text{eq}} = \frac{1}{3} \sum_i \sum_j U_{ij} a_i^* a_j^* a_i \cdot a_j$$

	x	y	z	U_{eq}
P	0.7918 (2)	0.2453 (1)	0.1453 (1)	0.0177 (6)
O1	0.7896 (5)	0.1255 (2)	0.1767 (2)	0.026 (2)
O2	0.9947 (5)	0.3039 (2)	0.2396 (3)	0.035 (2)
O3	0.7778 (4)	0.2604 (2)	0.0197 (2)	0.029 (2)
O4	0.5664 (5)	0.2987 (3)	0.1454 (3)	0.037 (3)
O5	0.4182 (5)	0.2209 (3)	0.3018 (3)	0.043 (3)
N1	0.9090 (6)	0.0815 (3)	0.4095 (3)	0.022 (3)
N2	0.9839 (7)	0.9607 (3)	0.1062 (3)	0.028 (3)
C1	0.8225 (8)	0.9225 (4)	-0.0127 (4)	0.028 (3)
C2	0.7699 (8)	-0.0141 (4)	0.4130 (4)	0.025 (4)
C3	1.2060 (9)	0.9927 (4)	0.1062 (4)	0.028 (3)
C4	1.1552 (7)	0.0542 (4)	0.4615 (4)	0.024 (3)

Table 2. Selected bond distances (\AA) and angles ($^\circ$) with *e.s.d.*'s in parentheses

P—O3	1.511 (2)	N2—C3	1.479 (7)
P—O2	1.519 (2)	C1—C3 ⁱⁱ	1.50 (1)
P—O1	1.524 (2)	O1—H2 ⁱⁱⁱ	1.89 (4)
P—O4	1.589 (3)	O1—H11	1.92 (6)
N1—C2	1.487 (7)	O2—H13 ^{iv}	1.72 (4)
N1—C4	1.479 (6)	O2—H8 ^v	1.79 (4)
C2—C4 ⁱ	1.489 (5)	O3—H14 ^{vi}	1.80 (3)
N2—C1	1.471 (9)	O3—H5 ^{vi}	1.89 (6)
O2—P—O3	112.70 (3)	O1—P—O2	112.20 (3)
O1—P—O3	111.80 (3)	O4—P—O2	107.50 (3)
O4—P—O3	106.00 (3)	O1—P—O4	107.30 (3)

Symmetry code: (i) $2-x, -y, 1-z$; (ii) $2-x, 2-y, -z$; (iii) $x, y-1, z$; (iv) $1+x, y, z$; (v) $2-x, y-\frac{1}{2}, \frac{1}{2}-z$; (vi) $x, \frac{1}{2}-y, z-\frac{1}{2}$.

In order to synthesize a fluoro-aluminophosphate with a microporous framework, $(\text{N}_2\text{C}_4\text{H}_{12})\cdot\text{HPO}_4\cdot\text{H}_2\text{O}$ was prepared hydrothermally from a 1:2:2:5:80 mixture of Al_2O_3 , P_2O_5 , NH_4F , piperazine and water. The resultant product was heated at 443 K for 3 d in an autoclave under autogenous pressure then filtered, washed and finally dried in air. A fragment of a colorless crystal without definite faces was isolated and its quality tested with Laue photographs.

Lists of structure factors, anisotropic thermal parameters, H-atom coordinates and complete geometry have been deposited with the British Library Document Supply Centre as Supplementary Publication No. SUP 71053 (10 pp.). Copies may be obtained through The Technical Editor, International Union of Crystallography, 5 Abbey Square, Chester CH1 2HU, England. [CIF reference: DU1011]

The authors thank Professor M. Leblanc and Dr R. Reoux (Université du Maine) for their help in data collection.

References

- Averbuch-Pouchot, M. T. & Durif, A. (1987). *Acta Cryst.* **C43**, 1894–1896.
 Averbuch-Pouchot, M. T., Durif, A. & Guitel, J. C. (1987). *Acta Cryst.* **C43**, 1896–1898.
 Cromer, D. T. & Mann, J. B. (1968). *Acta Cryst.* **A24**, 321–324.
 Sheldrick, G. M. (1976). *SHELX76*. Program for crystal structure determination. Univ. of Cambridge, England.
 Sheldrick, G. M. (1985). *SHELXS86*. In *Crystallographic Computing 3*, edited by G. M. Sheldrick, C. Krüger & R. Goddard, pp. 175–189. Oxford Univ. Press.
 Tapp, N. J., Milestone, N. B. & Bibby, D. M. (1988). *Zeolites*, Vol. 8, pp. 183–188.

Acta Cryst. (1993). **C49**, 1238–1240

(4*S*,5*S*)-4-Methyl-5-phenyl-1,3-oxazolidin-2-one

PAUL D. ROBINSON*

Department of Geology, Southern Illinois University, Carbondale, IL 62901, USA

DUY H. HUA,† LINDA A. GOOD AND HUI WANG

Department of Chemistry, Kansas State University, Manhattan, KS 66506, USA

CAL Y. MEYERS

Department of Chemistry and Biochemistry, Southern Illinois University, Carbondale, IL 62901, USA

(Received 15 September 1992; accepted 22 December 1992)

Abstract

Ring closure of methyl (2*S*,3*R*)-*N*-[(3-hydroxy-3-phenyl)-2-propyl]carbamate by treatment with *p*-toluene-sulfonyl chloride and pyridine leads to the title compound in which the oxazolidinone ring is puckered with C5 out of the plane by 21.56° .

Comment

In our studies of the stereoselective conversion of (-)-(1*R*,2*S*)-norephedrine into (+)-(1*S*,2*S*)-2-amino-1-phenylpropanol, we found that (2*S*,3*R*)-*N*-[(3-hydroxy-

† Fellow of the Alfred P. Sloan Foundation, 1989–1993.